

# Ionic compounds

5.2.1.3 · Chemistry · revision notes

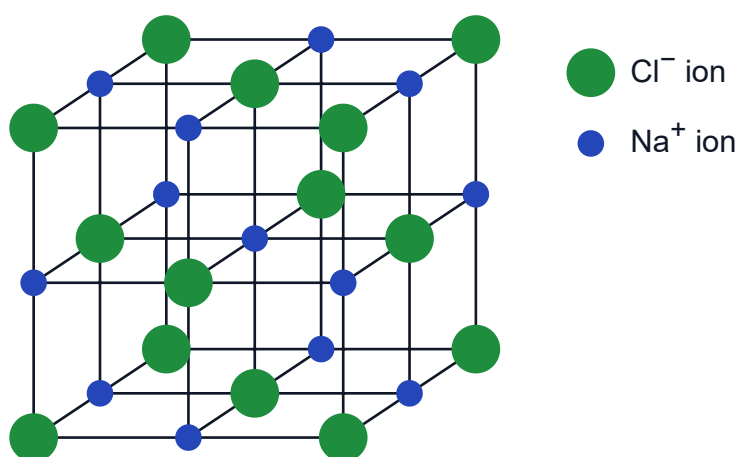
What an ionic compound is like: a giant lattice of oppositely charged ions held by strong electrostatic forces acting in all directions. You need to recognise ionic structures in diagrams, work out empirical formulae from diagrams, and give the limitations of each type of diagram.

## Key facts

- Ionic compound = giant lattice of oppositely charged ions
- No molecules: the formula gives the ratio of the ions
- Ball and stick: shows the 3D pattern, but the gaps and sticks are not real
- Every diagram shows only a tiny part of the lattice
- Held by strong electrostatic forces of attraction, acting in all directions
- Empirical formula = simplest whole-number ratio of ions
- 2D diagram: only one layer. Space-filling: can't see the ions inside

### Giant ionic lattices

- An ionic compound is a **giant structure of ions**, called a **giant ionic lattice**. The ions are arranged in a regular, repeating pattern.
- The ions are held together by **strong electrostatic forces of attraction** between **oppositely charged ions**. This is ionic bonding.
- The forces act in **all directions** in the lattice: each ion is attracted to all the oppositely charged ions around it, not just to one partner.
- In sodium chloride, each  $\text{Na}^+$  ion is surrounded by six  $\text{Cl}^-$  ions, and each  $\text{Cl}^-$  ion by six  $\text{Na}^+$  ions.



Ball and stick model of sodium chloride. The centre ion has six neighbours of the opposite charge.

- There are no molecules in an ionic compound. A formula such as  $\text{NaCl}$  gives the ratio of the ions.

### Diagrams of ionic structures and their limitations

- **Dot and cross diagram:** shows how the ions formed (electron transfer) and their charges. It does not show how the ions are arranged in the lattice or their relative sizes.
- **Ball and stick model:** shows the regular 3D arrangement of the ions. But the sticks are not real, there are no large gaps between ions in reality, and the forces act in all directions, not just along the sticks.
- **2D diagram:** easy to draw and shows the alternating pattern of ions, but only one layer, not the 3D arrangement.
- **3D space-filling diagram:** shows the ions touching and their relative sizes (a  $\text{Cl}^-$  ion is bigger than an  $\text{Na}^+$  ion). But you can only see the outside ions, and the forces are not shown.
- Every diagram shows only a tiny part of the lattice, which really continues in all directions.

### Empirical formula from a diagram

- The **empirical formula** is the simplest whole-number ratio of the ions (or atoms) in a compound.
- Count the ions of each type in the diagram, write the ratio, and cancel it down: 12 of X and 6 of Y gives X : Y = 2 : 1, so  $\text{X}_2\text{Y}$ .
- If charges are shown, check that they balance: two  $1+$  ions and one  $2-$  ion give zero overall.

## How to answer each type of question

### Describe the structure and bonding in an ionic compound

3 marks

Grade 5

1. Say it is a giant lattice (giant ionic structure).
2. Name the particles: oppositely charged ions, with their formulae if you can.
3. Say they are held by strong electrostatic forces of attraction acting in all directions.

**Example.** Describe the structure and bonding in magnesium oxide.

**Model answer.** Magnesium oxide is a giant (ionic) lattice (1) of  $\text{Mg}^{2+}$  ions and  $\text{O}^{2-}$  ions (1), held together by strong electrostatic forces of attraction between the oppositely charged ions, acting in all directions (1).

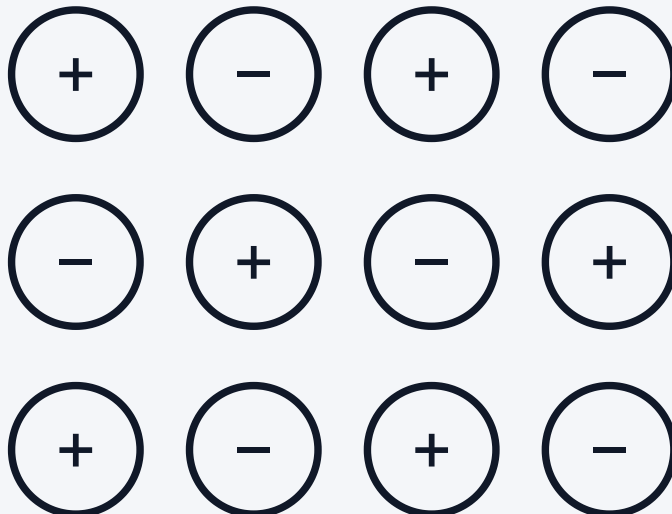
### Identify an ionic structure from a diagram

1 to 2 marks

Grade 5

1. Look for two types of particle, labelled + and – or with charges.
2. Check they are in a regular arrangement that continues in all directions, with no separate molecules.
3. Give the structure and the evidence you used.

**Example.** The diagram shows part of the structure of substance Z. The pattern continues in all directions.



What type of structure does Z have? Give a reason.

**Model answer.** Giant ionic lattice (1). It is made of oppositely charged ions in a regular pattern (1).

### Don't lose marks

- ✗ Calling an ionic compound 'molecules' or mentioning intermolecular forces. It is a giant lattice of ions.
- ✗ Saying the forces act only between pairs of ions or in one direction. They act in all directions.

- ✗ Writing the numbers counted (e.g.  $M_8Q_{16}$ ) instead of the simplest ratio.
- ✗ Giving a vague limitation such as 'it is not accurate'. Say exactly what is wrong, e.g. 'the ions are not really joined by sticks'.
- ✗ Writing 'strong forces' without saying they are electrostatic and between oppositely charged ions.

## More tips

### Memory tricks

- Picture a 3D chessboard of + and – ions that keeps going in every direction.
- Limitations checklist: sticks, gaps, scale, 3D, tiny part. Pick the ones that apply to the diagram in the question.
- Cancel a ratio of ions like a fraction:  $6 : 12 \rightarrow 1 : 2$ .

### Exam technique

- Learn the key phrase word for word: 'strong electrostatic forces of attraction between oppositely charged ions'.
- For advantage and disadvantage questions, make each point specific to the type of diagram named.
- If the diagram labels the ions with charges, use the charges to check your empirical formula.

## What each grade needs

What you need to be able to do, from the first marks up to the top grade.

### Grade 3

#### Describe an ionic compound as a giant structure

It is a giant structure (lattice) of ions arranged in a regular pattern.

### Grade 4

#### Name the force holding the ions together

Strong electrostatic forces of attraction between oppositely charged ions.

### Grade 5

#### State that the forces act in all directions

Each ion is attracted to all the oppositely charged ions around it, throughout the lattice.

### Grade 5

#### Recognise an ionic structure from a diagram

Look for a regular lattice of two types of ion, often labelled with charges, and no separate molecules.

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