

Group 7

5.1.2.6 · Chemistry · revision notes

The halogens in Group 7: their properties and trends, the compounds they form with metals and non-metals, why they get less reactive down the group, and displacement reactions. Displacement and 'explain the trend' questions are very common.

Key facts

- Group 7 = halogens: seven outer electrons; molecules of two atoms (Cl_2)
- Down the group: reactivity decreases
- With metals: ionic halides ($1-$ ions); with non-metals: covalent molecules
- $\text{Cl}_2 + 2\text{KBr} \rightarrow 2\text{KCl} + \text{Br}_2$ (colourless $\rightarrow$ orange)
- Down the group: relative molecular mass, melting point and boiling point increase
- Reason: outer shell further from nucleus $\rightarrow$ weaker attraction $\rightarrow$ harder to gain an electron
- A more reactive halogen displaces a less reactive one from a solution of its salt

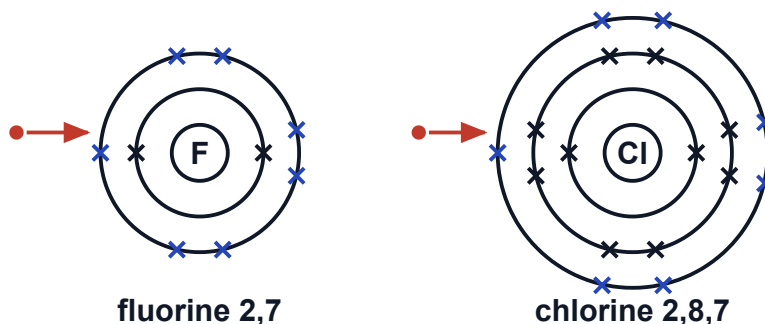
Notes

The halogens

- Group 7 contains fluorine (F), chlorine (Cl), bromine (Br), iodine (I) and astatine (At).
- They are **non-metals** and consist of **molecules made of pairs of atoms**: F_2 , Cl_2 , Br_2 , I_2 .
- They all have **seven electrons in their outer shell**, so they have similar reactions. In reactions with metals each atom gains one electron to form a halide ion with a $1-$ charge, e.g. Cl^- .
- At room temperature: chlorine is a pale green gas, bromine is a red-brown liquid and iodine is a grey-black solid.

Trends down the group

- Going down the group, the **relative molecular mass**, **melting point** and **boiling point** all increase.
- Going down the group, **reactivity decreases**.
- Why: the atoms have more shells, so the outer shell is further from the nucleus. The attraction for an extra electron is weaker, so it is harder for the atom to gain an electron.



the outer shell of F is nearer the nucleus, so it attracts an extra electron more strongly: F is more reactive

Each halogen atom gains one electron (red) to fill its outer shell.

Compounds of the halogens

- With **metals**, halogens form **ionic compounds** (metal halides) containing halide ions with a 1– charge: $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$.
- With **non-metals**, halogens form **covalent compounds** made of molecules, by sharing electrons: $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$ (hydrogen chloride).
- Names: chlorine → chloride, bromine → bromide, iodine → iodide.

Displacement reactions

- A more reactive halogen **displaces** a less reactive halogen from an aqueous solution of its salt.
- chlorine + potassium bromide → potassium chloride + bromine: $\text{Cl}_2 + 2\text{KBr} \rightarrow 2\text{KCl} + \text{Br}_2$. The colourless solution turns orange as bromine forms.
- Chlorine or bromine added to potassium iodide solution turns it brown as iodine forms.
- There is no reaction if the halogen added is less reactive than the one in the salt, e.g. iodine and potassium bromide.

How to answer each type of question

Predict and explain a displacement result

2 to 3 marks

Grade 5

1. Decide which halogen is more reactive (the one higher up the group).
2. A reaction happens only if the added halogen is more reactive than the one in the salt.
3. Give the colour change: orange for bromine, brown for iodine.

Example. A student adds bromine water to potassium chloride solution and to potassium iodide solution.

- (a) Which solution reacts? Explain your answer.
- (b) What colour does the solution that reacts turn?

Model answer. (a) Potassium iodide (1). Bromine is more reactive than iodine, so it displaces it, but bromine is less reactive than chlorine (1).

(b) Brown (1)

Describe the compounds halogens form

2 marks

Grade 5

1. With a metal: ionic compound, containing halide ions (1-).
2. With a non-metal: covalent compound, made of molecules.

Example. Chlorine reacts with sodium and with hydrogen.
Compare the compounds formed in these two reactions.

Model answer. With sodium, chlorine forms an ionic compound, sodium chloride, which contains Cl^- ions (1). With hydrogen, chlorine forms a covalent compound, hydrogen chloride, made of molecules (1).

Don't lose marks

- ✗ Naming the ion or compound with '-ine': NaBr is sodium bromide, not sodium bromine.
- ✗ Writing Cl instead of Cl_2 for chlorine.
- ✗ Saying halogen atoms lose electrons. They gain one electron.
- ✗ Explaining the trend as for Group 1 ('the electron is lost more easily'). For the halogens it is about gaining an electron.
- ✗ Saying iodine displaces chlorine or bromine.

More tips

Memory tricks

- Group 7 is the opposite of Group 1: the halogens get **less** reactive down the group.
- A halogen can displace only the halogens below it in the group.
- The colours get darker down the group: pale yellow (F_2), pale green (Cl_2), red-brown (Br_2), grey-black (I_2).
- Halogen to halide: change '-ine' to '-ide'.

Exam technique

- In displacement questions, first decide which halogen is more reactive, then say whether a reaction happens.
- Explain the trend in three steps: distance of the outer shell → strength of attraction → how easily an electron is gained.
- For a colour change, give the colour the solution turns (orange or brown), not the colour of the halogen you added.

What each grade needs

What you need to be able to do, from the first marks up to the top grade.

Grade 3

Recall the halogens and their states

At room temperature chlorine is a green gas, bromine a red-brown liquid and iodine a grey-black solid.

Grade 4

Describe the trends down Group 7

Relative molecular mass, melting point and boiling point increase; reactivity decreases.

Grade 5

Describe the compounds halogens form

Ionic compounds with metals (halide ions, $1-$); covalent molecules with non-metals, e.g. HCl .

Grade 5

Predict whether a displacement reaction happens

A more reactive halogen displaces a less reactive one from a solution of its salt.

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